



The lead is oxidized to PbO_2 at the anode.

The mass of lead lactate (FM 385.3) giving 0.111 1 g of PbO_2

(FM = 239.2) is $(385.3/239.2)(0.111 1 \text{ g}) = 0.179 0 \text{ g}$.

$$\% \text{ Pb} = \frac{0.179 0}{0.326 8} \times 100 = 54.77\%$$

